

Thermodynamics

Problem 1 (20%)

Describe briefly the following energy systems: (a) the combined cycle (b) fuel cell, (c) batteries, and (d) the cogeneration.

Problem 2 (20%)

The power required to operate a steady-state compressor is 3.56 Kw. Air enters the compressor at 1 bar and 300 K at a rate of 1 kg/min and leaves at 7 bars and 500 K. Determine (a) the rate of heat transfer, in kJ/h, (b) the entropy change of the air, in KJ/(min K), and (c) the entropy change of the environment which receives the heat transferred at 288 K, in kJ/(min K). (d) Is the process reversible, irreversible, or impossible?

problem 3 (20%)

At what temperature, will CO be 10% of the total moles of products if CO is burned with the stoichiometric amount of O<sub>2</sub> at 2-atm pressure?

Problem 4 (20%)

(a) Derive the expression  $c_p - c_v = -T(\partial v/\partial T)_p^2 (\partial P/\partial v)_T$ . (b) At 500 K the values of  $v$ ,  $\beta_p$ , and  $K_T$  for solid copper are 7.115 mL/gmol,  $54.2 \times 10^{-6} \text{ K}^{-1}$ , and  $0.837 \times 10^{-7} \text{ cm}^2/\text{N}$ , respectively. Determine the value of  $c_p - c_v$  in J/gmol C. ( $\beta = (1/v)(\partial v/\partial T)_p$ ,  $K_T = -(1/v)(\partial v/\partial P)_T$ ).

Problem 5 (20%)

A Carnot heat engine receive 90 kJ from a reservoir at 900 K. It rejects heat to the environment at 300 K. One-fifth of its work output is used to derive a Carnot refrigerator. The refrigerator rejects 60 kJ to the environment at 300 K. Find (a) the work output of the heat engine, (b) the efficiency of the heat engine, (c) the temperature of the low-temperature reservoir for the refrigerator, and (d) the coefficient of performance (COP) of the refrigerator.

Physical constants and conversion factors

Physical constants	
Avogadro's number	$N_A = 6.023 \times 10^{23}$ atoms/kgmol
Universal gas constant	$R_u = 0.08205$ L·atm/(kgmol·K) $= 8.314$ kJ/(kgmol·K) $= 0.08314$ bar·m <sup>3</sup> /(kgmol·K) $= 8.314$ kPa·m <sup>3</sup> /(kgmol·K)
Planck's constant	$h = 6.626 \times 10^{-34}$ J·s/molecule
Boltzmann's constant	$k = 1.380 \times 10^{-23}$ J/(K·molecule)
Speed of light	$c = 2.988 \times 10^{10}$ cm/s
Standard gravity	$g = 9.80665$ m/s <sup>2</sup>

Conversion factors	
1 cm	$= 0.3937$ in $= 10^4 \mu\text{m} = 10^8 \text{Å}$
1 km	$= 0.6215$ mi $= 3281$ ft
1 kg	$= 2.205$ lb <sub>m</sub>
1 N	$= 1 \text{ kg} \cdot \text{m/s}^2 = 0.2248$ lb <sub>f</sub>
1 bar	$= 10^5 \text{ N/m}^2 = 0.9869$ atm $= 100$ kPa
1 torr	$= 1 \text{ mmHg at } 0^\circ\text{C} = 1.333$ mbar $= 1.933 \times 10^{-2}$ psi
1 mbar	$= 0.402$ inHg
1 L	$= 0.0353$ ft <sup>3</sup> $= 0.2642$ gal $= 61.025$ in <sup>3</sup> $= 10^{-3}$ m <sup>3</sup>
1 g/cm <sup>3</sup>	$= 1 \text{ kg/L} = 62.4$ lb <sub>m</sub> /ft <sup>3</sup> $= 10^3$ kg/m <sup>3</sup>
1 J	$= 1 \text{ N} \cdot \text{m} = 1 \text{ V} \cdot \text{C}$ $= 0.7375$ ft·lb <sub>f</sub> $= 10$ bar·cm <sup>3</sup> $= 0.624 \times 10^6$ eV
1 kJ	$= 0.948$ Btu $= 737.6$ ft·lb <sub>f</sub> $= 10^{-2}$ bar·m <sup>3</sup>
1 kJ/kg	$= 0.431$ Btu/lb
1 W	$= 1 \text{ J/s} = 3.413$ Btu/h
1 kW	$= 1.3405$ hp $= 737.3$ ft·lb <sub>f</sub> /s
1 m/s	$= 2.237$ mi/h $= 3.60$ km/h $= 3.281$ ft/s
1 kJ/(kg·K)	$= 0.2389$ Btu/(lb <sub>m</sub> ·°F)
$T(\text{K})$	$= \frac{5}{9}(T(^\circ\text{F}) + 459.67) = T(^\circ\text{C}) + 273.15 = T(^{\circ}\text{R})/1.8$

Derived units and common multipliers

1. Some SI derived units

Physical quantity	Unit	Symbol	Definition
force	newton	N	1 kg·m/s <sup>2</sup>
pressure	pascal	Pa	1 kg/m·s <sup>2</sup> (= 1 N/m <sup>2</sup> )
pressure	bar	bar	10 <sup>5</sup> kg/m·s <sup>2</sup> (= 10 <sup>5</sup> N/m <sup>2</sup> )
energy	joule	J	1 kg·m <sup>2</sup> /s <sup>2</sup> (= 1 N·m)
power	watt	W	1 kg·m <sup>2</sup> /s <sup>3</sup> (= 1 J/s)
electric quantity	coulomb	C	1 A·s
electric potential	volt	V	1 kg·m <sup>2</sup> /(A·s <sup>2</sup> ) (= 1 A·Ω)
resistance	ohm	Ω	1 kg·m <sup>2</sup> /(A <sup>2</sup> ·s <sup>2</sup> ) (= 1 V/A)
capacitance	farad	F	1 A <sup>2</sup> ·s <sup>4</sup> /(kg·m <sup>2</sup> ) (= 1 C/V)

2. Names and symbols for common multipliers of SI units

Multiplier	Prefix	Symbol
	giga	G
	mega	M
	kilo	k
	deci	d
	centi	c
	milli	m
	micro	μ
	nano	n

Logarithms to the base 10 of the equilibrium constant  $K_p$

$$K_p = \frac{(p_E)^{v_E} (p_F)^{v_F}}{(p_A)^{v_A} (p_B)^{v_B}}$$
 for the reaction  $v_A A + v_B B \rightleftharpoons v_E E + v_F F$

Numbered reactions:

- (1)  $\text{H}_2 \rightleftharpoons 2\text{H}$
- (2)  $\text{O}_2 \rightleftharpoons 2\text{O}$
- (3)  $\text{N}_2 \rightleftharpoons 2\text{N}$
- (4)  $\text{O}_2 + \text{N}_2 \rightleftharpoons \text{NO}$
- (5)  $\text{H}_2\text{O} \rightleftharpoons \text{H}_2 + \frac{1}{2}\text{O}_2$
- (6)  $\text{H}_2\text{O} \rightleftharpoons \text{OH} + \frac{1}{2}\text{H}_2$
- (7)  $\text{CO}_2 \rightleftharpoons \text{CO} + \frac{1}{2}\text{O}_2$
- (8)  $\text{CO}_2 + \text{H}_2 \rightleftharpoons \text{CO} + \text{H}_2\text{O}$

Temp., K	log $K_p$ values for reactions numbered above							
	(1)	(2)	(3)	(4)	(5)	(6)	(7)	(8)
298	-71.224	-81.208	-159.600	-15.171	-40.048	-46.054	-45.066	-5.018
500	-40.316	-45.880	-92.672	-8.783	-22.886	-26.130	-25.025	-2.139
1000	-17.292	-19.614	-43.056	-4.062	-10.062	-11.280	-10.221	-0.159
1200	-13.414	-15.208	-34.754	-3.275	-7.899	-8.789	-7.764	+0.135
1400	-10.630	-12.054	-28.812	-2.712	-6.347	-7.003	-6.014	+0.333
1600	-8.532	-9.684	-24.350	-2.290	-5.180	-5.662	-4.706	+0.474
1700	-7.666	-8.706	-22.512	-2.116	-4.699	-5.109	-4.169	+0.530
1800	-6.896	-7.836	-20.874	-1.962	-4.270	-4.617	-3.693	+0.577
1900	-6.204	-7.058	-19.410	-1.823	-3.886	-4.177	-3.267	+0.619
2000	-5.580	-6.356	-18.092	-1.699	-3.540	-3.780	-2.884	+0.656
2100	-5.016	-5.720	-16.898	-1.586	-3.227	-3.422	-2.539	+0.688
2200	-4.502	-5.142	-15.810	-1.484	-2.942	-3.095	-2.226	+0.716
2300	-4.032	-4.614	-14.818	-1.391	-2.682	-2.798	-1.940	+0.742
2400	-3.600	-4.130	-13.908	-1.305	-2.443	-2.525	-1.679	+0.764
2500	-3.202	-3.684	-13.070	-1.227	-2.224	-2.274	-1.440	+0.784
2600	-2.836	-3.272	-12.298	-1.154	-2.021	-2.042	-1.219	+0.802
2700	-2.494	-2.892	-11.580	-1.087	-1.833	-1.828	-1.015	+0.818
2800	-2.178	-2.536	-10.914	-1.025	-1.658	-1.628	-0.825	+0.833
2900	-1.882	-2.206	-10.294	-0.967	-1.495	-1.442	-0.649	+0.846
3000	-1.606	-1.898	-9.716	-0.913	-1.343	-1.269	-0.485	+0.858
3100	-1.348	-1.610	-9.174	-0.863	-1.201	-1.107	-0.332	+0.869
3200	-1.106	-1.340	-8.664	-0.815	-1.067	-0.955	-0.189	+0.878
3300	-0.878	-1.086	-8.186	-0.771	-0.942	-0.813	-0.054	+0.888
3400	-0.664	-0.846	-7.736	-0.729	-0.824	-0.679	+0.071	+0.895
3500	-0.462	-0.620	-7.312	-0.690	-0.712	-0.552	+0.190	+0.902

Source: Based on data from the JANAF Tables, NSRDS-NBS-37, 1971, and revisions published in *Journal of Physical and Chemical Data* through 1982.

Ideal-gas properties of air

T, K; h, kJ/kg; u, kJ/kg; s\*, kJ/(kg·K)

T	h	u	s*	$\ln p_r$	$\ln v_r$	$\ln p_r$	T	h	u	s*
200	199.97	0.3363	1.29559	6.245	329.97	2.13407	460	462.02	142.56	1.7071
210	209.97	0.3987	1.34444	6.742	337.32	2.15604	470	472.34	149.69	1.7152
220	219.97	0.4690	1.39105	7.242	344.70	2.17760	480	482.49	156.82	1.7233
230	230.02	0.5477	1.43557	7.724	352.08	2.19876	490	492.74	164.00	1.7314
240	240.02	0.6355	1.47824	8.191	359.49	2.21952	500	503.02	171.13	1.7395
250	250.05	0.7329	1.51917	8.643	366.92	2.23993	510	513.32	178.28	1.7476
260	260.09	0.8405	1.55848	9.081	374.36	2.25997	520	523.63	185.45	1.7557
270	270.11	0.9590	1.59634	9.505	381.84	2.27967	530	533.95	192.60	1.7638
280	280.13	1.0889	1.63279	9.916	389.34	2.29906	540	544.25	199.75	1.7719
285	285.14	1.1584	1.65055	10.116	396.86	2.31809	550	554.54	206.91	1.7800
290	290.16	1.2311	1.66802	10.303	404.42	2.33685	560	564.82	214.07	1.7881
295	295.17	1.3068	1.68515	10.477	411.97	2.35531	570	575.09	221.24	1.7962
300	300.19	1.3860	1.70203	10.648	419.55	2.37348	580	585.34	228.42	1.8043
305	305.22	1.4686	1.71865	10.816	427.15	2.39140	590	595.52	235.61	1.8124
310	310.24	1.5546	1.73498	10.981	434.78	2.40902	600	605.70	242.82	1.8205
315	315.27	1.6442	1.75106	11.143	442.42	2.42644	610	615.87	250.05	1.8286
320	320.29	1.7375	1.76690	11.302	450.09	2.44356	620	626.03	257.24	1.8367
325	325.31	1.8345	1.78249	11.458	457.78	2.46048	630	636.18	264.46	1.8448
330	330.34	1.9352	1.79783	11.612	465.50	2.47716	640	646.32	271.69	1.8529
340	340.42	2.149	1.82790	11.764	473.25	2.49364	650	656.45	278.93	1.8610
350	350.49	2.379	1.85706	11.913	481.01	2.50985	660	666.57	286.16	1.8691
360	360.58	2.626	1.88543	12.059	488.81	2.52589	670	676.68	293.43	1.8772
370	370.67	2.892	1.91313	12.202	496.62	2.54175	680	686.78	300.69	1.8853
380	380.77	3.176	1.94001	12.342	504.45	2.55731	690	696.87	307.99	1.8934
390	390.88	3.481	1.96633	12.479	512.30	2.57277	700	706.95	315.30	1.9015
400	400.98	3.806	1.99194	12.613	520.23	2.58810	710	717.02	322.62	1.9096
410	411.12	4.153	2.01699	12.744	528.14	2.60319	720	727.08	330.00	1.9177
420	421.26	4.522	2.04142	12.872	536.07	2.61803	730	737.13	337.43	1.9258
430	431.43	4.915	2.06533	12.997	544.02	2.63270	740	747.17	344.91	1.9339
440	441.61	5.332	2.08870	13.119	551.99	2.64737	750	757.20	352.43	1.9420
450	451.80	5.775	2.11161	13.239	560.01	2.66176	760	767.22	360.00	1.9501